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Chapter 12 REVIEW:

Stoichiometry,

Theoretical, Actual &

Percent yield Part I.

Stoichiometry 1. 1 ZnI_2

$\rightarrow 1 \text{ Zn} + 1 \text{ I}_2$ How

many grams of iodine

will you produce if you

begin your rxn with

56.7g of zinc iodide?

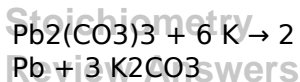
56.7 g 1 mole 1 mole

253.8 g 319.19 g 1

mole 1 mole = 45.08

grams of iodine 2. 1

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Chapter 12 REVIEW: Stoichiometry, Theoretical, Actual

...

Chapter 12 Review
Stoichiometry

Important Vocabulary •

Stoichiometry- The calculations of quantities in chemical reactions in a subject of chemistry • Mole ratio- a conversion factor derived from the

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coefficients of a
balanced chemical
equation interpreted in
terms of moles

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Stoichiometry Section Review In Example

12.2.1 and Example 12.2.2, the identity of the limiting reactant has been apparent: $[\text{Au}(\text{CN})_2]^-$, LaCl_3 , ethanol, and para-nitrophenol. When the limiting reactant is not apparent, we can determine which

Chapter 12 **Stoichiometry** **Section Review** **Answer Key**

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Stoichiometry.
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Stoichiometry.
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Important Vocabulary.

- Stoichiometry- The calculations of quantities in chemical reactions in a subject of chemistry
- Mole ratio- a conversion factor derived from the coefficients of a balanced chemical equation interpreted in terms of moles
- Limiting reagent- The reactant that determines the amount

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of product that can be formed by a reaction • Excess reagent- The reaction that is not completely used up in a reaction • Theoretical yield- Maximum ...

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Stoichiometry.
Stoichiometry. Is the number of atoms conserved in a b.... Is

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the number of molecules conserved in.... Is the number of moles conserved in a b.... The calculation of quantities in chemical reactions. Yes... $N_2 + 3H_2 \rightarrow 2NH_3$... 2 atoms + 6 atoms \rightarrow 8 atoms.

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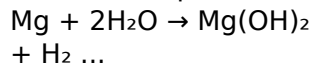
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Stoichiometry. Terms
in this set (21)

Stoichiometry. The
calculation of
quantities in chemical
reactions is a subject of
chemistry. Mole ratio.

... Use the following
balanced equation to
answer the question:



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Chapter Test A

Answer Key

Stoichiometry 379

CHAPTER 12

Assessment 36. a. Two formula units KClO_3 decompose to form two formula units KCl and three molecules O_2 . b. Four molecules NH_3 react with six molecules NO to form five molecules N_2 and six molecules H_2O . c. Four atoms K react

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with one molecule O_2
to form two formula
units K_2O . 37. a. Two
mol $KClO_3$ decompose
to

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Chapter 12 **Stoichiometry** **Section Review** **Answer Key**

Chapter 12

Stoichiometry 12.1 The
Arithmetic of Equations

12.2 Chemical

Calculations 12.3

Limiting Reagent and
Percent Yield . . .

excess of 1800 is a

logical answer. • The

unit of the known (FSW

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3 HP 2) cancels. • The answer has the correct unit (W). 3 Evaluate Does the result make sense?

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Chemistry Chapter
12: Stoichiometry ...**

Chapter 12 Multiple
Choice Identify the
letter of the choice that
best completes the
statement or answers
the question. ____ 1.
The first step in most
stoichiometry problems
is to ____ . a. add the
coefficients of the
reagents c. convert
given quantities to

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volumes b. convert
given quantities to
moles d. convert given
quantities to masses
____ 2.

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Prioritize Schedule your
time realisti-cally

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Arithmetic of Equations

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