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This equation can be used to calculate the pH of a solution containing the acid and one of its salts, that is, of a buffer solution. With bases, if the value of an equilibrium constant is known in the form of a base association constant, K_b the dissociation constant of the conjugate acid may be calculated from $pK_a + pK_b = 14$...

Henderson-Hasselba
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Hasselbalch Equation Its History And **Ich equation - Wikipedia**

The Henderson-Hasselbalch equation provides a relationship between the pH of acids (in aqueous solutions) and their pK_a (acid dissociation constant). The pH of a buffer solution can be estimated with the help of this equation when the concentration of the acid and its conjugate base, or the base and the

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corresponding
conjugate acid, are
known.

Henderson- Hasselbalch Equation - Estimating the pH of Buffers

Example Problem

Applying the
Henderson-Hasselbalch
Equation . Calculate
the pH of a buffer
solution made from
0.20 M $\text{HC}_2\text{H}_3\text{O}_2$
and 0.50 M $\text{C}_2\text{H}_3\text{O}_2^-$

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2-that has an acid dissociation constant for $\text{HC}_2\text{H}_3\text{O}_2$ of 1.8×10^{-5} .

Henderson- Hasselbalch Equation and Example

The Henderson-Hasselbalch equation was developed independently by the American biological chemist L. J. Henderson and the Swedish

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physiologist K. A.
Hasselbalch, for
relating the pH to the
bicarbonate buffer
system of the blood
(see below). In its
general form, the
Henderson-Hasselbalch
equation is a useful
expression for buffer
calculations.

Henderson- Hasselbalch Equation - an overview ...

However, the

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Henderson-Hasselbalch equation for a basic buffer will give pOH and its pH can be calculated as $(14 - \text{pOH})$. 2. The dissociation constant of a weak acid (or weak base) can be determined by measuring the pH of a buffer solution containing equimolar concentrations of the acid (or base) and the salt.

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equation and its
Significance

The Henderson-Hasselbalch equation provides a general solution to the quantitative treatment of acid-base equilibrium in a biological system. This article explains how to derive the Henderson-Hasselbalch Equation. Buffers are the mixture of weak acids and their

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salts of strong bases
(or) the mixture of
weak bases and their
salts of strong
acids. Simply, Buffers
are an important
concept of Acid ...

What is Henderson- Hasselbalch Equation? (Basic Concept)

Many students of
chemistry have
wondered if putting the
mass action expression
in logarithmic format

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should have warranted immortalization of the names Henderson and Hasselbalch. With focus on this question, this article examines the evolution of the Henderson-Hasselbalch equation and presents a critical evaluation of its usefulness.

The Henderson- Hasselbalch Equation: Its History and ...

What is its pH?

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Solution: The Henderson-Hasselbalch Equation (done in the Internet way): $\text{pH} = \text{pK}_a + \log [\text{base} / \text{acid}]$
 $x = 4.752 + \log (0.800 / 1.00)$
 $x = 4.752 - 0.097 = 4.655$. Note how decreasing the amount of base makes the buffer pH become more acidic (compare to example #1).

**ChemTeam: Buffers
and the Henderson-
Hasselbalch**

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The Henderson-Hasselbalch equation can be also used in the case of polyprotic acids, as long as the consecutive pK_a values differ by at least 2 (better 3). Thus it can be safely used in the case of phosphoric buffers ($pK_{a1} = 2.148$, $pK_{a2} = 7.199$, $pK_{a3} = 12.35$), but not in the case of citric acid ($pK_{a1} = 3.128$, $pK_{a2} = 4.761$, pK_{a3}

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(=6.396).

Henderson- Hasselbalch

equation and the pH of buffer solution

“The Henderson-
Hasselbalch Equation:
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Wasserstoffzahl des
Blutes aus der freien
und gebundenen
Kohlensäure

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desselben, und die
 Sauerstoffbindung des
 Blutes als Funktion der
 Wasserstoffzahl."

Biochemische
 Zeitschrift, 1917,
 pp.112-144.

pH, pKa, and the He nderson-Hasselbalch Equation

The
 Henderson-Hasselbalc
 h equation
 mathematically
 connects the
 measurable pH of a

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solution with the pK_a (which is equal to $-\log K_a$) of the acid. The equation is also useful for estimating the pH of a buffer solution and finding the equilibrium pH in an acid-base reaction.

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(Equation $\left(\text{ref}\{HH\}\right)$)
is an approximation,
with a certain region of
validity. By its nature,
it does not take into
account the self-
dissociation of water,
which becomes
increasingly important
in dilute solutions.
When concentrations
reach somewhere
around $\left(10^{-5}\right)\text{ mol}\text{L}^{-1}$
or lower, the
true $\left(\text{pH} \dots\right)$

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Answer to: Using the Henderson-Hasselbalch equation, calculate the pH of a buffer solution that is 0.320 M in HCO_3^- and 0.354 M in CO_3^{2-} . By...

Using the Henderson-Hasselbalch equation, calculate the pH ...

Hasselbalch equation and comments on its

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Hasselbalch history, including the development of the pH scale. Henderson Equation We will discuss the limitations of the Henderson-Hasselbalch equation focusing on the titration curve of a weak acid with a strong base. Over much of the titration range, the calculation of pH relies on the Henderson-Hasselbalch ...

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The Henderson-Hasselbalch Equation: Its History and ...

The Henderson-Hasselbalch equation is useful in estimating the pH of buffer solutions as can be seen in the applications and examples page. Check Out: pKa table and its use; Applications and example problems based on Henderson-Hasselbalch equation

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Equation:

Derivation |

Analytical ...

The Henderson-Hasselbalch equation is a tool that allows us to calculate the pH of an acid solution using the pKa of the acid and the relative concentrations of the acid and its conjugate base. It is defined as:

**Henderson-
Hasselbalch
Equation - MCAT**

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The Henderson-Hasselbalch equation is widely used by many scientists especially chemists, biologists and pharmacists. We have already seen the derivation of the Henderson-Hasselbalch equation. On this page we discuss some of the applications of the Henderson-Hasselbalch equation and solve some sample problems which one may

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Applications and
Example Problems
Using Henderson ...

encounter in research.

Applications and Example Problems Using Henderson ...

Henderson-Hasselbalch equation is a simple expression which relates the pH, pKa and the buffer action of a weak acid and its conjugate base. The Henderson-Hasselbalch equation also describes the characteristic shape of the titration curve of any weak acid

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such as acetic acid,
phosphoric acid, or any
amino acid.

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00998ecf8427e.](#)